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(54) Process for producing particles and coating films of titanium oxide.

(57) The present invention relates to a process for producing titanium oxide particles and a process for producing a titanium oxide coating film. The titanium oxide particles are produced by subjecting a suspension of titanium oxide fine particles to a hydrothermal treatment at 100 °C or higher at a pH of a specific range to carry out crystal growth of the fine particles. The titanium oxide coating film is produced by suspending the thus obtained titanium oxide particles in water, adjusting the suspension to a pH of a specific range, coating the suspension on a support and then firing the coat to allow the titanium oxide particles to firmly adhere to the support.

The titanium oxide particles and titanium oxide coating film obtained according to the present invention are used as pigments, ultraviolet absorbers, fillers, optical materials, electronic material, optoelectric transducing materials, decorative materials, catalysts, photocatalysts, catalyst carriers, adsorbents and bioreactors.

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## BACKGROUND OF THE INVENTION

## Field of the Invention

5 The present invention relates to a process for producing particles and coating films of titanium oxide and more particularly to a process for producing particles and coating films of titanium oxide which are used for pigments, ultraviolet absorbers, fillers, optical materials, electronic materials, optoelectric transducers, decorative materials, catalysts, photocatalysts, catalyst carriers, adsorbents, bioreactors and the like.

10 Description of Related Art

Titanium oxide particles are useful as pigments, ultraviolet absorbers, fillers and the like. Furthermore, titanium oxide coating films comprising titanium oxide particles allowed to firmly adhere to a support are  
 15 useful as photocatalysts, sensors and materials for semiconductor electrodes such as solar batteries utilizing the photoconductor characteristics of titanium oxide and as sealing materials utilizing its dielectric characteristics. Moreover, titanium oxide coating films are used as heat ray reflecting glasses and decorative materials because of high light reflectance and beautiful silvery reflective color tone and furthermore as catalysts, photocatalysts, catalyst carriers, adsorbents and bioreactors utilizing the porosity  
 20 of the titanium oxide coating films. The titanium oxide coating films can be produced, for example, by a process which comprises immersing a support in a solution containing titanium isopropoxide or the like, taking it out and then heating the support at a high temperature to form a film, a process which comprises blowing titanium tetrachloride vapor against the surface of a support at a high temperature to form a film, a process which comprises spraying a solution containing an organotitanium compound such as titanium acetylacetonato, titanium alkoxide or the like to the surface of a support at a high temperature to form a  
 25 film.

According to the above-mentioned conventional processes, a titanium compound is reacted on a support to form a titanium oxide film. Therefore, the size of titanium oxide primary particles which constitute the resulting titanium oxide film is several nm and such titanium oxide particles show volume shrinkage at  
 30 the time of sintering to cause occurrence of cracks in the resulting film. The film having cracks is apt to peel off from the support and is low in endurance and wear resistance and furthermore, inferior in optical characteristics. In addition, according to the conventional processes, it is difficult to make porous titanium oxide coating films having a large thickness and a large specific surface area which are suitable for optoelectric transducing materials, catalysts, photocatalysts and adsorbents. The conventional processes  
 35 have further problems that the working atmosphere becomes worse due to the vapors of organic solvents and titanium compounds and there are dangers in working such as firing.

## SUMMARY OF THE INVENTION

40 An object of the present invention is to provide titanium oxide particles excellent in crystallinity and extremely excellent in various properties. Another object of the present invention is to provide titanium oxide coating films less in cracks and high in adhesion to a support.

The present invention is a process for producing titanium oxide particles which comprises subjecting titanium oxide fine particles to a hydrothermal treatment at a temperature of 100°C or higher, characterized  
 45 by carrying out the hydrothermal treatment using (1) a suspension of titanium oxide fine particles containing rutile-type titanium oxide, (2) a suspension of titanium oxide fine particles containing anatase-type titanium oxide obtained by hydrolyzing titanyl sulfate or (3) a suspension in which various titanium oxide fine particles such as hydrous titanium oxide fine particles of metatitanic acid or orthotitanic acid in addition to the above-mentioned titanium oxide fine particles are highly dispersed. Furthermore, the present invention  
 50 is a process for producing a titanium oxide coating film on a support using the above titanium oxide particles.

## BRIEF DESCRIPTION OF THE DRAWINGS

55 Fig. 1 is a optical photomicrograph (x 1000) which shows the surface morphologic structure of the titanium oxide coating film obtained in Example 3.

Fig. 2 is an electron photomicrograph (x 8000) which shows the layer structure of a section of the titanium oxide coating film obtained in Example 7.

## DETAILED DESCRIPTION OF PREFERRED EMBODIMENTS

The inventors have conducted intensive research in an attempt to obtain a titanium oxide coating film strong in adhesion to a support and excellent in optical properties and as a result, have found that (1) by using titanium oxide particles obtained by carrying out crystal growth of titanium oxide fine particles, volume shrinkage at the time of formation of the film can be reduced, occurrence of cracks can be reduced and adhesion to the support can be improved, (2) the titanium oxide particles which are allowed to firmly adhere are desirably those which have an average particle size of 10 to 100 nm, preferably 10 to 30 nm, especially preferably 15 to 25 nm because these are excellent in adhesion to the support, (3) for carrying out the crystal growth of titanium oxide fine particles, there are the following methods of (a) heating titanium oxide fine particles at a temperature of 200 to 600 °C, (b) subjecting the titanium oxide fine particles to a hydrothermal treatment at a temperature of 100 °C or higher, and (c) subjecting a suspension in which the titanium oxide fine particles are highly dispersed to a hydrothermal treatment at a temperature of 100 °C or higher, and these methods are preferred because the crystals of the titanium oxide fine particles can be uniformly grown, and especially preferred is the method (b) in which titanium oxide fine particles containing rutile-type titanium oxide or titanium oxide fine particles obtained by hydrolyzing titanyl sulfate are subjected to a hydrothermal treatment, and (4) in the case of forming a titanium oxide coating film by allowing titanium oxide particles obtained by crystal growth of titanium oxide fine particles to firmly adhere to a support, the packing density of the titanium oxide coating film can be enhanced and adhesion to the support can be further improved by using a suspension in which said titanium oxide particles are highly dispersed. Thus, the present invention has been accomplished.

In the present invention, the "titanium oxide" means various titanium oxides such as anatase-type titanium oxide, rutile-type titanium oxide, amorphous titanium oxide, metatitanic acid, orthotitanic acid and the like and hydrous titanium oxide. Moreover, the "titanium oxide fine particles" means those which have a fine average particle size of 3 to 20 nm, preferably 5 to 15 nm and the "titanium oxide particles" means those having an average particle size of 10 to 100 nm, preferably 10 to 30 nm, especially preferably 15 to 25 nm obtained by crystal growth of said titanium oxide fine particles. The said titanium oxide fine particles are obtained by known processes, for example, by hydrolyzing or neutralizing titanium compounds such as titanyl sulfate, titanium chloride, titanium acetate, titanium alkoxide and the like, if necessary, in the presence of a seed. Then, the thus obtained titanium oxide fine particles are preferably filtrated and washed for removing impurities such as sulfates, chlorine and alkalis.

The "titanium oxide fine particles containing rutile-type titanium oxide" used herein means titanium oxide comprising 0.5 to 50% by weight, preferably 1 to 30% by weight of fine particles of titanium oxide which can be identified to be rutile-type crystal in the X-ray diffraction pattern and the balance being mostly amorphous. One example corresponding thereto is the rutile transition accelerating seed used in preparation of titanium dioxide pigments. The said titanium oxide fine particles containing rutile-type titanium oxide can be prepared, for example, by the following processes: ① A titanium compound such as titanyl sulfate, titanium acetate or titanium tetrachloride is neutralized with an alkali or hydrolyzed, if necessary, in the presence of nucleus crystals to obtain a precipitate and the resulting precipitate is heated, if necessary, with addition of an alkali and then is aged at a temperature of from 50 °C to the boiling point with addition of an inorganic acid such as hydrochloric acid or nitric acid or an organic acid such as citric acid. ② A Titanium chloride is neutralized with an alkali at a pH of 4 to 7. ③ Titanium chloride is hydrolyzed by heating under the condition of 11% or higher in hydrochloric acid concentration. ④ Titanium chloride is subjected to an oxidative thermal decomposition in vapor phase.

The "titanium oxide fine particles obtained by hydrolysis of titanyl sulfate" used herein means titanium oxide fine particles containing anatase-type titanium oxide obtained by hydrolyzing titanyl sulfate with heating at 60 to 110 °C, especially 80 to 110 °C. This contains a slight amount of amorphous titanium oxide, but contains mostly titanium oxide which can be identified to be anatase-type crystal in the X-ray diffraction pattern.

When the heating temperature at the time of hydrolysis is lower than 60 °C, hydrolysis of titanyl sulfate is not sufficiently performed or a long time is required for hydrolysis. The concentration of titanyl sulfate in the solution and the heating time can be optionally set, but a titanyl sulfate concentration of 50 to 250 g/l in terms of TiO<sub>2</sub> is suitable and a heating time of about 1 to 10 hours is suitable. The pressure at the time of heating is preferably atmospheric pressure, but the heating may be conducted under application of pressure. Furthermore, if necessary, a seed may be present in the solution of titanyl sulfate. Preferably, the thus obtained titanium oxide fine particles are filtrated and washed to remove impurities such as sulfate radicals.

The "suspension in which titanium oxide fine particles are highly dispersed" used herein is obtained by allowing an acid to be present in a suspension of titanium oxide fine particles to adjust the pH to 4 or lower, preferably 2 or lower or allowing an alkali to be present in a suspension of titanium oxide fine particles to adjust the pH to 9 or higher, preferably 10 or higher and dispersing titanium oxide fine particles or by grinding the powders of titanium oxide fine particles and then preparing a suspension thereof, or by suspending titanium oxide fine particles in a suspension with grinding the fine particles by a grinder or a classifier, or by allowing a dispersant to be present in the suspension.

When acid or alkali is allowed to be present, if pH of the suspension is higher than 4 and lower than 9, titanium oxide fine particles are dispersed with difficulty and cannot be made to a colloid solution. For adjusting the pH to 4 or lower, at least one acid optionally selected from inorganic acids such as hydrochloric acid, nitric acid, sulfuric acid, phosphoric acid and carbonic acid and organic acids such as citric acid, malic acid, acetic acid and lactic acid is added to the suspension. Hydrochloric acid or nitric acid is preferred. For adjusting the pH to 9 or higher, at least one alkali optionally selected from sodium hydroxide, potassium hydroxide, sodium carbonate, ammonium carbonate, ammonia and amines is allowed to be present. The grinders used are preferably wet grinders such as sand mill, ball mill, Henschel mixer and paint shaker. As the dispersant when this is used, mention may be made of, for example, inorganic dispersants such as sodium hexametaphosphate and sodium silicate and organic dispersants such as sodium polycarboxylates, polyethylene glycol and acetylacetone. These dispersants may be used each alone or in combination of two or more. In the present invention, organic dispersants are especially preferred and sodium polycarboxylates such as OROTAN 850 (manufactured by Rohm & Haas Co.) and polyethylene glycol are especially preferred among the organic dispersants. These dispersants can be added optionally to the suspension of titanium oxide fine particles before coating of the suspension on a support, thereby to allow them to be present in the suspension. The amount of the dispersant to be added is preferably about 0.1 to 40% by weight in terms of the effective ingredient based on the weight of the titanium oxide fine particles. In the present invention, preferred are the process of allowing an acid to be present in the suspension of titanium oxide fine particles to adjust the pH to 4 or lower, preferably 2 or lower or allowing an alkali to be present in the suspension of titanium oxide fine particles to adjust the pH to 9 or higher, preferably 10 or higher and dispersing the titanium oxide fine particles.

The hydrothermal treatment is carried out by heating the suspension or highly dispersed suspension of titanium oxide fine particles to a temperature of 100°C or higher, preferably 100 to 250°C. By this hydrothermal treatment, crystals of the titanium oxide fine particles are grown to an average particle size of 10 to 100 nm, preferably 10 to 30 nm, especially preferably 15 to 25 nm. Furthermore, by this hydrothermal treatment, the titanium oxide fine particles containing rutile-type titanium oxide become rutile-type titanium oxide particles and X-ray diffraction of this rutile-type titanium oxide particles reveals no anatase-type crystal and this rutile-type titanium oxide particles have substantially rutile-type crystal structure. The titanium oxide fine particles containing anatase-type titanium oxide become anatase-type titanium oxide particles by the hydrothermal treatment. X-ray diffraction of this anatase-type titanium oxide particles reveals that the particles show substantially anatase-type crystal structure. The thus obtained rutile-type titanium oxide particles or anatase-type titanium oxide particles can have the desired particle size by suitably setting the temperature and time of the hydrothermal treatment. By subjecting the suspension in which titanium oxide fine particles are highly dispersed to the hydrothermal treatment, the effect of the hydrothermal treatment can be further enhanced.

The hydrothermal treatment is preferably carried out under about saturated vapor pressure, but can also be carried out under higher than the saturated vapor pressure or under about atmospheric pressure and can normally be carried out by industrially employed heat-resisting and pressure-resisting apparatuses. The time for the hydrothermal treatment can be optionally set, but is suitably about 1 to 48 hours.

If necessary, the titanium oxide particles obtained by the hydrothermal treatment are filtrated and washed to remove impurities such as sulfate radicals, chlorine and alkalis, and dried to make dry powders. The drying can be carried out at optional temperatures, but suitable is 100 to 500°C. Furthermore, it is also possible depending on the using sites to use the titanium oxide particles by supporting or coating them on a support or by grinding or molding the above dry powders. It is also possible depending on the catalytic reactions used to support transition metals such as gold, platinum, rhodium and palladium on the surface of the titanium oxide particles.

The present invention further relates to a process for producing a titanium oxide coating film by coating on a support the titanium oxide particles obtained by carrying out the crystal growth of the titanium oxide fine particles and firing the coated particles. The "titanium oxide coating film" in the present invention means a coating film containing substantially no resin, being different from the general coating films comprising a titanium oxide pigment and a resin medium. Furthermore, it has a large specific surface area

of about 10 to 200 cm<sup>2</sup> per 1 cm<sup>2</sup> of the support and 1 μm of a film thickness. The titanium oxide coating film of the present invention can be optionally changed in its thickness, transparency and porosity expressed by specific surface area by suitably setting the conditions such as concentration of titanium oxide suspension and particle size of the titanium oxide particles explained hereinafter. Moreover, the titanium oxide coating film of either anatase-type or rutile-type can be formed by suitably selecting the crystal system of the titanium oxide particles used or by suitably setting the firing conditions. Furthermore, it is also possible to form a coating film comprising a mixture of anatase-type titanium oxide particles and rutile-type titanium oxide particles at an optional proportion. The thickness of the film can be, for example, about 0.5 to 20 μm. The titanium oxide particles having an average particle size of 10 to 100 nm, preferably 10 to 30 nm, especially preferably 15 to 25 nm are superior in adhesion to the support.

The method for growing the crystals of the titanium oxide fine particles includes, for example, (a) a method of heating them at a temperature of 200 to 600 °C, preferably 300 to 500 °C, especially preferably 300 to 450 °C and (b) the above-mentioned method of subjecting them to hydrothermal treatment at a temperature of 100 °C or higher, preferably 100 to 250 °C.

When the heating temperature is lower than 200 °C, crystal growth of the titanium oxide fine particles is not sufficient and volume shrinkage of titanium oxide particles when allowed to firmly adhere to a support is great and cracks are apt to occur and adhesion to the support becomes inferior. When the heating temperature is higher than 600 °C, the crystals of the titanium oxide fine particles grow too much and adhesion to the support is deteriorated or the specific surface area of the titanium oxide coating film decreases. The heating time is suitably 0.5 to 2 hours.

By carrying out the above heating treatment or hydrothermal treatment, crystals of the titanium oxide fine particles grow and crystallinity is improved and uniform titanium oxide particles of suitable size can be obtained. Thus, the desired titanium oxide coating film can be obtained.

In the present invention, the titanium oxide particles obtained by carrying out the crystal growth of the titanium oxide particles are first suspended in a solvent such as water, alcohol, toluene or the like. As the solvent, water is especially preferred from the point of working atmosphere. The concentration of the titanium oxide particles in the suspension can be optionally set, but a concentration of 10 to 1200 g/l in terms of TiO<sub>2</sub> is preferred. When the concentration is lower than 10 g/l, it becomes difficult to increase the thickness of the titanium oxide coating film. When it is higher than 1200 g/l, viscosity of the suspension increases and the suspension can be coated on a support with difficulty. In the present invention, it is preferred to increase the concentration by subjecting the titanium oxide particles in the suspension to the aforementioned dispersion treatment. Furthermore, the suspension of the titanium oxide particles obtained by the hydrothermal treatment can be coated as it is or after adjusting the concentration of the titanium oxide particles without carrying out the filtration of the titanium oxide particles on a support by various coating methods mentioned hereafter and thus, a titanium oxide coating film can be obtained.

The resulting suspension of titanium oxide particles is coated on a support by usual methods such as dip coating, spray coating, spinner coating, blade coating, roller coating, wire bar coating and reverse roll coating and, if necessary, dried and fired to remove the solvent, and the titanium oxide particles are allowed to firmly adhere to the support. Material, shape and size of the support on which the titanium oxide film is formed can be optionally selected depending on the object and its use. As the material there may be used glass, metals, ceramics and others, and as the shape and size mention may be made of a card board, a small fragment, spheres such as beads. The firing is carried out at 100 °C or higher, preferably 200 to 800 °C, especially preferably 300 to 800 °C. When the firing temperature is lower than 100 °C, adhesion to the support deteriorates. The upper limit of the firing temperature depends on the use and cannot be generically specified, but the firing temperature higher than 800 °C is not preferred in some case because crystal growth of the titanium oxide particles is too much or the specific surface area of the titanium oxide coating film is small. In the case of carrying out the heating for crystal growth, the firing is carried out at a temperature higher than the heating temperature. When the firing temperature for adhesion of the particles is lower than the heating temperature for the crystal growth, adhesion to the support is inferior.

In the present invention, it is preferred to use titanium oxide particles obtained by subjecting the suspension in which the titanium oxide fine particles are highly dispersed to the hydrothermal treatment. Furthermore, for producing a titanium oxide coating film comprising rutile-type titanium oxide particles, it is preferred to use the titanium oxide particles comprising substantially rutile-type titanium oxide obtained by subjecting the titanium oxide fine particles containing rutile-type titanium oxide to a hydrothermal treatment at a temperature of 100 °C or higher. For producing a titanium oxide coating film of anatase-type titanium oxide, it is preferred to use the substantially anatase-type titanium oxide particles obtained by hydrolyzing titanyl sulfate and then subjecting the resulting titanium oxide fine particles to a hydrothermal treatment at a temperature of 100 °C or higher.

In this way, the titanium oxide film of the present invention is obtained.  
The present invention is illustrated by the following examples.

#### Example 1

One liter of an aqueous solution containing 200 g/l of titanium tetrachloride was added dropwise to 2 liters of an aqueous solution containing 222 g/l of sodium carbonate with stirring to neutralize titanium tetrachloride at room temperature. Subsequently, to the above aqueous solution was added citric acid in an amount of 5% by weight based on the weight of  $\text{TiO}_2$  of the neutralization precipitate. Then, the aqueous solution was heated to  $70^\circ\text{C}$  and aged for 20 minutes. (A part of the resulting titanium oxide fine particles were taken, filtrated and dried. X-ray diffraction thereof showed that they contained 15% by weight of titanium oxide fine particles having rutile-type crystal structure, the balance being amorphous.)

The thus obtained titanium oxide fine particles were filtrated, washed and then dispersed in water to prepare a suspension of 250 g/l in terms of  $\text{TiO}_2$ . Then, aqueous nitric acid solution was added to this suspension to adjust to a pH of 1.3. Thereafter, the suspension was charged in an autoclave and subjected to a hydrothermal treatment at  $150^\circ\text{C}$  for 13 hours under a saturated vapor pressure. The resulting product was filtrated, washed and dried to obtain the titanium oxide particles of the present invention (Sample A). X-ray diffraction of this Sample A showed that it had substantially the rutile-type crystal structure and had a particle size of 18 nm. Furthermore, it had a cubic form which is euhedral for the rutile-type crystal according to observation under an electron microscope.

#### Comparative Example 1

Aqueous ammonia was added to 1 liter of a solution containing 80 g/l of titanium tetrachloride with stirring to adjust the pH to 7.0 to obtain a gel. This gel was filtrated, washed and then dispersed in water to prepare a suspension of 250 g/l in terms of  $\text{TiO}_2$ . This suspension was charged in an autoclave and subjected to a hydrothermal treatment at  $150^\circ\text{C}$  for 13 hours under a saturated vapor pressure. The resulting product was filtrated, washed and dried to obtain titanium oxide particles (Sample B). X-ray diffraction of this Sample B showed that it had an anatase-type crystal structure and had a particle size of 10.5 nm.

#### Comparative Example 2

The commercially available titanium oxide catalyst P-25 (a mixture of rutile-type and anatase-type, particle size: 22 nm, manufactured by Japan Aerosil Co.) was employed as Sample C.

The catalytic activity of Samples A-C obtained in Example 1 and Comparative Examples 1 and 2 was examined in the following manner. Each of the samples was dispersed in water to prepare a suspension of 4 g/l in terms of  $\text{TiO}_2$ . To 25 ml of the resulting suspension was added 10  $\mu\text{l}$  of tetrachloroethylene and then, the suspension was irradiated with a light from a xenon lamp of 150 W for 30 minutes to carry out a photocatalytic decomposition reaction of tetrachloroethylene. The decomposition rate of tetrachloroethylene with the sample was calculated from the concentration of tetrachloroethylene before the reaction and that of tetrachloroethylene after the reaction. The results are shown in Table 1. As can be seen from the results, the rutile-type titanium oxide particles of the present invention are superior in the catalytic activity.

Table 1

	Sample	Decomposition rate (%)
Example	A	72
Comparative Example 1	B	42
Comparative Example 2	C	48

Furthermore, crystallinity of Samples A-C was examined by spectroscopic means using a photo-acoustic spectroscopic method. That is, when 0.5 g of the sample is charged in a closed container and irradiated with a light of a given wavelength, the sample absorbs the light depending on the wavelength. In the case of emitting the energy of this absorbed light as a heat by nonradiative process, the atmosphere

around the sample vibrates and generates a sound wave due to the thermal expansion of the sample with increase of the vibration energy of crystal lattice. The sound wave generated in this case is detected by a high-sensitivity microphone. The wavelength of the light with which the sample is irradiated is changed from 300 to 1600 nm and the output of the high-sensitivity microphone is measured to obtain a photo-acoustic spectrum of each sample.

It is generally said that rutile-type titanium oxide has an energy gap of 3.0 eV. When an energy corresponding to this energy gap is converted into a light energy, this is a light of 413 nm in wavelength. Therefore, theoretically, when a light slightly shorter in wavelength than 413 nm is irradiated, the light is completely absorbed by titanium oxide and when a light slightly longer than 413 nm is irradiated, the light is utterly not absorbed by titanium oxide. However, titanium oxide inferior in crystallinity has a defect level capable of absorbing an energy smaller than 3.0 eV and absorbs even the light slightly longer than 413 nm. From this fact, superiority or inferiority of crystallinity of the samples can be determined by comparing the state of this specific light absorption edge using the above-mentioned photo-acoustic spectrum. From the fact that the photo-acoustic spectrum of the rutile-type titanium oxide particles of the present invention was sharp in rising of the specific light absorption edge, it was seen that the rutile-type titanium oxide particles of the present invention were superior in crystallinity.

#### Example 2

One liter of a solution containing 80 g/l of titanyl sulfate was heated to 85°C and kept for 3 hours to hydrolyze titanyl sulfate. (A part of the resulting titanium oxide fine particles were taken, filtrated and dried. X-ray diffraction of the product showed that it had an anatase-type crystal structure and had a particle size of 5 nm.)

The thus obtained titanium oxide fine particles were filtrated, washed and then dispersed in water to prepare a suspension of 250 g/l in terms of  $\text{TiO}_2$ . Then, aqueous hydrochloric acid solution was added to this suspension to adjust the pH to 1.0. Thereafter, the suspension was charged in an autoclave and subjected to a hydrothermal treatment at 150°C for 13 hours under a saturated vapor pressure. The resulting product was filtrated, washed and dried to obtain the titanium oxide particles of the present invention (Sample D). X-ray diffraction of this Sample D showed that it had an anatase-type crystal structure and had a particle size of 13 nm.

#### Comparative Example 3

Titanium oxide fine particles were obtained in the same manner as in Example 2 except that the hydrothermal treatment was not conducted (Sample E). X-ray diffraction of this Sample E showed that it had an anatase-type crystal structure and a particle size of 5 nm.

#### Comparative Example 4

The titanium oxide fine particles obtained in Comparative Example 3 were fired at 450°C for 30 minutes to obtain Sample F. X-ray diffraction of this Sample F showed that it had an anatase-type crystal structure and a particle size of 12 nm.

The catalytic activity of Samples B-F obtained in Example 2 and Comparative Examples 1-4 was examined in the following manner. Each of the samples was dispersed in water to prepare a suspension of 4 g/l in terms of  $\text{TiO}_2$ . To 25 ml of the resulting suspension was added 25  $\mu\text{l}$  of 2-propanol and then, the suspension was irradiated with a light from a xenon lamp of 150 W for 2 hours to carry out a photocatalytic decomposition reaction of 2-propanol. The decomposition rate of 2-propanol with the sample was calculated from the concentrations of 2-propanol before and after the reaction. The results are shown in Table 2. As can be seen from the results, the anatase-type titanium oxide particles of the present invention are superior in the catalytic activity.

Table 2

	Sample	Decomposition Rate (R)	
		One hour after starting of the reaction	2 hours after starting of the reaction
Example 2	D	33.5	54.7
Comparative Example 1	B	11.3	24.4
Comparative Example 2	C	24.6	40.8
Comparative Example 3	E	12.2	26.0
Comparative Example 4	F	22.5	35.5

Furthermore, the catalytic activity for the decomposition reaction of tetrachloroethylene was examined in the same manner as above. The results are shown in Table 3. As can be seen from the results, the anatase-type titanium oxide particles of the present invention are superior in catalytic activity.

Table 3

	Sample	Decomposition Rate (R)	
		15 minutes after starting of the reaction	30 minutes after starting of the reaction
Example 2	D	41.9	71.3
Comparative Example 2	C	28.8	47.5

The crystallinity of the Samples B-F was examined as mentioned hereinbefore. From the fact that the photo-acoustic spectrum of the anatase-type titanium oxide particles of the present invention was sharp in rising of the specific light absorption edge, it was seen that the anatase-type titanium oxide particles of the present invention were superior in crystallinity.

As mentioned above, the rutile-type titanium oxide particles and the anatase-type titanium oxide particles of the present invention are higher in catalytic activity than conventional ones. The reason is not clear, but can be considered as follows: The rutile-type titanium oxide particles and the anatase-type titanium oxide particles of the present invention are superior in crystallinity and a few lattice defects are present inside the titanium oxide particles. Since conventional rutile-type titanium oxide particles and anatase-type titanium oxide particles have many lattice defects, many electrons or holes generated by irradiation with ultraviolet rays or the like retain in the lattice defects and only a few electrons or holes can participate in the catalytic reaction. On the other hand, since the rutile-type titanium oxide particles and anatase-type titanium oxide particles of the present invention have only a few lattice defects, the electrons and holes generated hardly retain in the lattice defects and in addition, the charge separation between the electrons and holes generated becomes easy and the amount of electrons or holes capable of participating in the catalytic reaction increases. Furthermore, since the anatase-type titanium oxide particles of the present invention are obtained by subjecting the titanium oxide fine particles obtained by hydrolyzing titanyl sulfate to a hydrothermal treatment, many acid points are present on the surface of the resulting anatase-type titanium oxide particles and these acid points act as adsorption active points for the reactants to further enhance the catalytic activity.

### Example 3

One liter of a solution containing 80 g/l of titanyl sulfate was heated to 85°C to carry out hydrolysis. Then, the product was filtrated and washed to obtain a wet cake. Successively, water in this wet cake was replaced with acetone and then, the cake was subjected to vacuum drying at 60°C to obtain titanium oxide fine particle powders. (X-ray diffraction of the resulting titanium oxide fine particles showed that the product had anatase-type crystal structure and had a particle size of 5 nm.)



Then, the titanium oxide fine particle powders were heated at 300°C for 1 hour to obtain titanium oxide particles having anatase-type crystal structure of 11 nm in particle size. Then, the resulting titanium oxide particles were ground by an attritor and introduced into water. Thereafter, to this water was added polyethylene glycol (average molecular weight: 15000-25000) in an amount of 20% by weight of TiO<sub>2</sub>, followed by sufficient shaking by a paint conditioner (manufactured by Red Devil Co.) to obtain a suspension containing 500 g/l of titanium oxide particles. Then, this suspension was coated on a glass plate by a doctor blade (slit: 60 μm) and air-dried and then, fired at 450°C for 2 hours to obtain a titanium oxide coating film (anatase-type crystal, hazed) of the present invention.

This titanium oxide coating film firmly adhered to the glass plate (support) and the film was observed under a light microscope to find substantially no cracks on the surface. This titanium oxide coating film had a thickness of 10 μm measured by a feeler type film thickness gauge and had a specific surface area of 80 to 100 cm<sup>2</sup> per 1 cm<sup>2</sup> of the support and 1 μm of the film thickness which was measured by nitrogen gas adsorption method.

#### Example 4

A titanium oxide coating film (anatase-type crystal, hazed) of the present invention was obtained in the same manner as in Example 3 except that the calcination temperature was 400°C.

The resulting titanium oxide coating film firmly adhered to the glass plate (support) and the film was observed under a light microscope to find substantially no cracks. This titanium oxide coating film had a thickness of 10 μm and had a specific surface area of 80 to 100 cm<sup>2</sup> per 1 cm<sup>2</sup> of the support and 1 μm of the film thickness.

#### Example 5

An acid titania sol (CS-N, titanium oxide fine particles of anatase-type crystal and 5 nm in particle size, manufactured by Ishihara Sangyo Kaisha, Ltd.) was dried at 110°C for 5 hours. The resulting titanium oxide fine particle powders were heated at 300°C for 1 hour in the air to obtain titanium oxide particle powders (anatase-type crystal, particle size 11 nm). Then, the following components were shaken in a paint shaker for 5 hours to sufficiently mix and disperse them. This was coated on a glass plate of 100 cm<sup>2</sup> by a doctor blade and dried and then, was fired at 450°C for 30 minutes to obtain a titanium oxide coating film (anatase-type crystal, translucent) of the present invention.

The resulting titanium oxide coating film firmly adhered to the glass plate (support) and the film was observed under a light microscope to find substantially no cracks. This titanium oxide coating film had a thickness of 2.0 μm.

Titanium oxide particle powders	8.0 g
Polyethylene glycol (molecular weight 20000)	3.2 g
Pure water	14.0 ml
Concentrated nitric acid	0.5 ml

Catalytic activity of the resulting titanium oxide coating film was examined in the following manner. The titanium oxide coating film was put in a glass container of 8 liters and acetaldehyde which was a malodorous component was added thereto at a concentration of 50 ppm. Then, the glass container was sealed. The titanium oxide coating film was irradiated with a black light for 2 hours so that ultraviolet intensity on the surface of the film reached 1 mW/cm<sup>2</sup> and the concentration of acetaldehyde after the irradiation was measured to obtain 15 ppm. From this result, it can be seen that the titanium oxide coating film of the present invention is superior in photocatalytic activity.

#### Comparative Example 5

Water was added to the wet cake which was obtained in the same manner as in Example 3, but was not subjected to the heating and furthermore, polyethylene glycol (average molecular weight: 15000-25000) was added thereto in an amount of 20% by weight based on the titanium oxide, followed by stirring to obtain a suspension containing 300 g/l of titanium oxide fine particles.

Then, this suspension was coated on a glass plate by a doctor blade and air-dried. Thereafter, the coat was fired at 450°C for 2 hours. As a result, the titanium oxide fine particles peeled off from the glass plate

in the form of fragments and a titanium oxide coating film was not able to be formed on the support.

#### Example 6

5 One liter of a solution containing 80 g/l of titanyl sulfate was heated to 85°C and kept at that temperature for 3 hours to hydrolyze titanyl sulfate to obtain titanium oxide fine particles. (A part of the resulting titanium oxide fine particles were taken and filtrated, washed and dried. X-ray diffraction of the resulting product showed that the product had anatase-type crystal structure and had a particle size of 6 nm.)

10 The resulting titanium oxide fine particles were filtrated, washed and then dispersed in water to prepare a suspension of 200 g/l in terms of  $\text{TiO}_2$ . Then, aqueous hydrochloric acid solution was added to this suspension to adjust the pH to 1.0. Thereafter, the suspension was charged in an autoclave and subjected to a hydrothermal treatment at 200°C for 13 hours under a saturated vapor pressure. (A part of the resulting suspension of titanium oxide particles was taken and filtrated, washed and dried. X-ray diffraction of the resulting product showed that the product had anatase-type crystal structure and had a particle size of 20 nm.)

Then, the suspension of titanium oxide particles obtained by the hydrothermal treatment was coated on a glass plate by a doctor blade and air-dried. Thereafter, the coat was fired at 450°C for 3 hours to obtain a titanium oxide coating film (anatase-type crystal, transparent) of the present invention.

20 This titanium oxide coating film firmly adhered to the glass plate (support). The film was observed under a light microscope to find substantially no cracks. This titanium oxide coating film had a thickness of 6  $\mu\text{m}$  and had a specific surface area of 80 to 100  $\text{cm}^2$  per 1  $\text{cm}^2$  of the support and 1  $\mu\text{m}$  of the film thickness.

#### 25 Example 7

Titanyl sulfate was hydrolyzed in accordance with the process mentioned in Example 6 to obtain titanium oxide fine particles. (A part of the resulting titanium oxide fine particles were taken and filtrated, washed and dried. X-ray diffraction of the resulting product showed that it had anatase-type crystal structure and had a particle size of 5 nm.)

30 The resulting titanium oxide fine particles were filtrated, washed and then dispersed in water to prepare a suspension of 250 g/l in terms of  $\text{TiO}_2$ . Then, aqueous nitric acid solution was added to this suspension to adjust the pH to 1.2. Thereafter, the suspension was charged in an autoclave and subjected to a hydrothermal treatment at 150°C for 13 hours under a saturated vapor pressure. (A part of the resulting suspension of titanium oxide particles was taken and filtrated, washed and dried. X-ray diffraction of the resulting product showed that it had anatase-type crystal structure and had a particle size of 13 nm.)

To the suspension of the titanium oxide particles obtained by the above hydrothermal treatment was added polyethylene glycol (average molecular weight: 20000) in an amount of 40% by weight based on  $\text{TiO}_2$  in the suspension, followed by heating to 60°C. Then, in the same manner as in Example 6, the suspension was coated on a glass plate, air-dried, and then fired to obtain a titanium oxide coating film (anatase-type crystal, transparent) of the present invention.

40 This titanium oxide coating film firmly adhered to the glass plate (support). The film was observed under a light microscope to find substantially no cracks. This titanium oxide coating film had a thickness of 7.5  $\mu\text{m}$  and had a specific surface area of 100 to 120  $\text{cm}^2$  per 1  $\text{cm}^2$  of the support and 1  $\mu\text{m}$  of the film thickness.

#### Example 8

50 Water was added to the suspension of titanium oxide particles obtained in accordance with the process mentioned in Example 7 to adjust the concentration to 12 g/l based on  $\text{TiO}_2$ . To 200 ml of this suspension were added 10 g of commercially available hollow glass beads (B2B/750 manufactured by 3M Co.) and then, water was removed by an evaporator. The resulting product was dried at 100°C for 2 hours and then, fired at 300°C for 1 hour to obtain a titanium oxide coating film (anatase-type crystal, transparent) of the present invention.

55 This titanium oxide coating film firmly adhered to the hollow glass beads (support) and substantially no cracks were found according to the observation by a light microscope.

## Example 9

A titanium oxide coating film (anatase-type crystal, transparent) of the present invention was obtained in the same manner as in Example 8 except that solid glass beads (GB503M manufactured by Toshiba Balotyni Co.) were used in place of the hollow glass beads.

This titanium oxide coating film firmly adhered to the solid glass beads (support) and substantially no cracks were found according to the observation by a light microscope.

## Example 10

A titanium oxide coating film (anatase-type crystal, transparent) of the present invention was obtained in the same manner as in Example 8 except that a ceramic honeycomb structure (HONEYCERAM manufactured by Nihon Gaishi K.K.) was used in place of the hollow glass beads.

This titanium oxide coating film firmly adhered to the honeycomb structure (support) and substantially no cracks were found according to the observation by a light microscope.

## Example 11

Sodium hydroxide was added to an acid titania sol (CS-N, titanium oxide fine particles of anatase-type crystal having a particle size of 6 nm, manufactured by Ishihara Sangyo Kaisha, Ltd.) to adjust the pH to 7, followed by carrying out filtration and washing. Water was added to the resulting wet cake to prepare a suspension of titanium oxide fine particles having a concentration of 100 g/l in terms of  $\text{TiO}_2$ . Then, sodium hydroxide was added to the suspension to adjust the pH to 10 and the suspension was charged in an autoclave and subjected to a hydrothermal treatment at 150°C for 3 hours under a saturated vapor pressure.

Then, nitric acid was added to the suspension of titanium oxide particles (anatase-type crystal, particle size 18 nm) after subjected to the hydrothermal treatment to neutralize to a pH of 7, followed by carrying out filtration and water washing. Water was further added to the resulting cake of titanium oxide particles to prepare a suspension of titanium oxide particles having a concentration of 500 g/l in terms of  $\text{TiO}_2$ . Nitric acid was added to the suspension to adjust the pH to 1. This suspension was coated on a glass plate of 100  $\text{cm}^2$  by a spin coater and dried and then, fired at 500°C for 3 hours to obtain a titanium oxide coating film (anatase-type crystal structure, translucent).

This titanium oxide coating film firmly adhered to the glass plate (support) and substantially no cracks were found according to the observation by a light microscope. This titanium oxide coating film of the present invention had a thickness of 1.0  $\mu\text{m}$ .

The catalytic activity of the resulting titanium oxide coating film was examined by the method used in Example 5. As a result, the acetaldehyde concentration after irradiated with light was 13 ppm. It can be seen from the result that the titanium oxide coating film of the present invention is superior in photocatalytic activity.

The process of the present invention is an industrially very useful process according to which titanium oxide particles excellent in crystallinity and furthermore, in optical properties, optoelectric transducing properties and catalytic properties can be produced simply and easily. Furthermore, a titanium oxide coating film having a large specific surface area, less in occurrence of cracks and good in adhesion to a support can be produced by allowing the titanium oxide particles of the present invention to firmly adhere to a support.

The titanium oxide coating film of the present invention is also excellent in optical properties, optoelectric transducing properties and catalytic properties and is useful for optical materials, electronic materials, optoelectric transducing materials, decorative materials, catalysts, photocatalysts, catalyst carriers, adsorbents and bioreactors. The titanium oxide coating film of the present invention is especially excellent in photocatalytic activity, and injurious materials can be rapidly and efficiently removed utilizing the excellent photocatalytic activity. Thus, it is markedly useful not only for industrial uses, but also as household deodorants and the like. In addition, the titanium oxide coating film of the present invention can be used for a long period of time, is high in safety, can be applied to removal of a wide variety of injurious materials and does not pollute the environment even if it is abandoned. Therefore, it is industrially very useful.

Moreover, according to the process of the present invention, the titanium oxide coating film can be simply allowed to adhere to supports such as glass, metals, ceramics and plastics and further, the resulting film has a stable quality.

Claims

1. A process for producing a titanium oxide coating film which comprises carrying out crystal growth of titanium oxide fine particles to prepare titanium oxide particles, coating the resulting titanium oxide particles on a support and firing the coated support to allow the particles to firmly adhere to the support.
2. A process according to claim 1 wherein the titanium oxide particles have an average particle size of 10 to 100 nm.
3. A process according to claim 1 which comprises carrying out the crystal growth of the titanium oxide fine particles by heating the fine particles at a temperature of 200 to 600 °C to prepare titanium oxide particles, then coating the resulting titanium oxide particles on a support and firing the coated support at a temperature higher than the heating temperature for the crystal growth to allow the particles to firmly adhere to the support.
4. A process according to claim 1 wherein the titanium oxide fine particles is subjected to a hydrothermal treatment at 100 °C or higher to carry out the crystal growth of the fine particles.
5. A process according to claim 1 wherein a suspension of the titanium oxide fine particles is adjusted to a pH of 4 or lower or 9 or higher and then, the suspension is subjected to a hydrothermal treatment at 100 °C or higher to carry out the crystal growth of the titanium oxide fine particles.
6. A process according to claim 1 wherein the titanium oxide fine particles containing rutile-type titanium oxide is subjected to a hydrothermal treatment at 100 °C or higher to carry out the crystal growth of the fine particles.
7. A process according to claim 1 wherein titanyl sulfate is hydrolyzed to obtain the titanium oxide fine particles and the resulting titanium oxide fine particles are subjected to a hydrothermal treatment at 100 °C or higher to carry out the crystal growth of the fine particles.
8. A process according to claim 1 wherein a suspension of the titanium oxide particles is adjusted to a pH of 4 or lower or 9 or higher and then the suspension is coated on a support and fired to allow the titanium oxide particles to firmly adhere to the support.
9. A process for producing titanium oxide particles which comprises adjusting a suspension of titanium oxide fine particles to a pH of 4 or lower or 9 or higher and then, subjecting the suspension to a hydrothermal treatment at 100 °C or higher to carry out crystal growth of the fine particles.
10. A process for producing titanium oxide particles which comprises subjecting titanium oxide fine particles containing rutile-type titanium oxide to a hydrothermal treatment at 100 °C or higher to carry out crystal growth of the fine particles.
11. A process for producing titanium oxide particles which comprises hydrolyzing titanyl sulfate to obtain titanium oxide fine particles and subjecting the resulting titanium oxide fine particles to a hydrothermal treatment at 100 °C or higher to carry out crystal growth of the fine particles.

FIG. 1

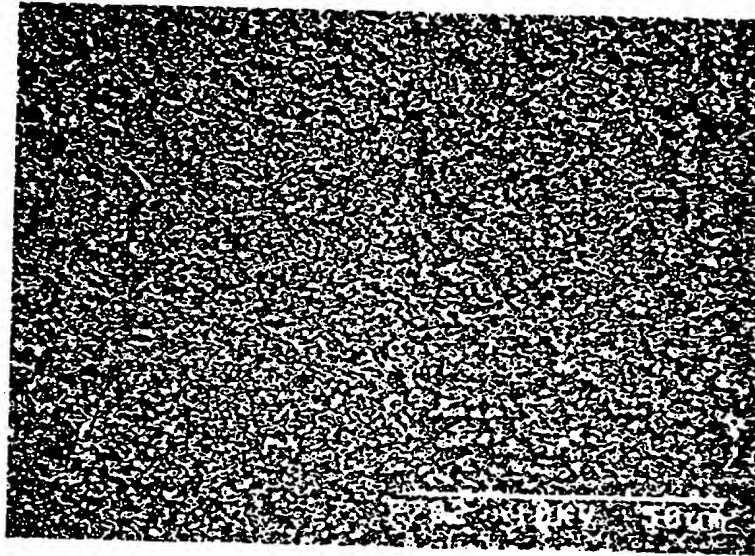
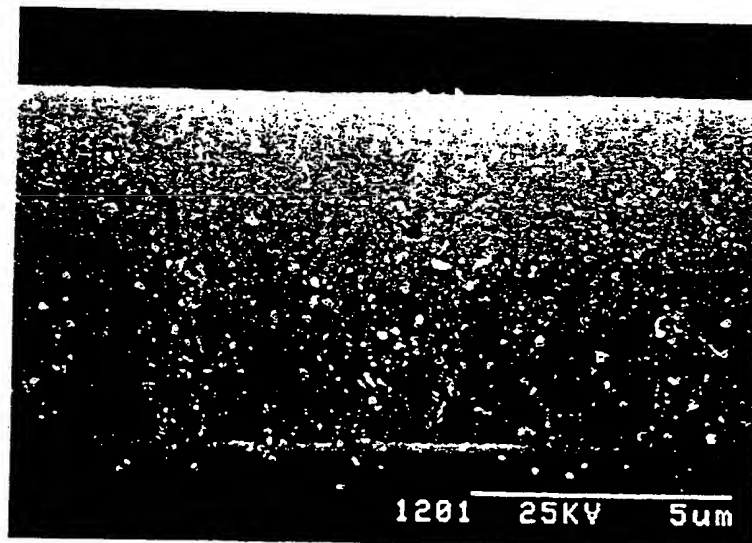


FIG. 2





European Patent  
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# EUROPEAN SEARCH REPORT

Application Number

EP 93 11 1854

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int. Cl.5)
A	GB-A-2 018 234 (CENTRAL GLASS COMPANY) * claim 1 *	1	C03C17/25 C01G23/053
A	CERAMICS INTERNATIONAL vol. 15, no. 5, 1989, pages 255 - 270 POUSKOULELI * page 264 *	9-11	
A	DATABASE WPI Week 8637, Derwent Publications Ltd., London, GB; AN 86-240930 & JP-A-61 168 528 (SHOKUBAI KASEI KOGYO) * abstract *	1	
A	APPLIED OPTICS vol. 26, no. 21, November 1987, pages 4688 - 4691 THOMAS	1	
			TECHNICAL FIELDS SEARCHED (Int. Cl.5)
			C03C C01G
The present search report has been drawn up for all claims			
Place of search BERLIN		Date of completion of the search 08 NOVEMBER 1993	Examiner CLEMENT J-P.
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